therefore suggest that the lengthening of the C(1)-C(7)bond is caused by this rotation, which has decreased the overlap between the two  $p_z$  orbitals and therefore decreased the double-bond character of the bond.

The molecule is distorted from planarity by an unusual amount. Omitting the 'outside' atoms of the isobutyl group, C(10), C(11), C(12), the distortion of one half of the molecule from complete planarity may be approximately described by a rotation of the (nearly) planar group C(1), C(7), C(8), N, C(9), Cu through an angle of  $25^{\circ}$  about the line C(1)-C(9). Equations for various least-squares planes and the distances of the several atoms from these planes are given in Table 4. Comparison of these distances with those found in other compounds shows a number of differences. The metal atom is farther (1.12 Å) out of the plane defined by the benzene ring, the farthest heretofore reported being 0.75 Å in bis-(N-t-butylsalicylaldiminato)palladium(II) (Day, Glick & Hoard, 1968). The nitrogen atom is also farther (0.25 Å) out of the plane and is out in the same direction as the metal, whereas in other compounds it is out in the opposite direction from the metal atom (cf. Table 6, Jain & Lingafelter, 1967). The dihedral angle between the coordination plane  $CuN_2O_2$ and the mean plane of the o-hydroxyacetophenone imine residue is 136.6°, which may be compared with the corresponding angle of 144.4° in bis-(N-t-butylsalicylaldiminato)palladium.

The packing of the molecules may be seen in Fig. 2. None of the intermolecular contact distances is unusual.

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# The Crystal Structure of Compounds with (N-P)<sub>n</sub> Rings. VII.\* Refinement of the Crystal Structure of Hexabromocyclotriphosphazene, N<sub>3</sub>P<sub>3</sub>Br<sub>6</sub>

# By H. ZOER AND A. J. WAGNER

Laboratory of Structural Chemistry, The University, Bloemsingel 10, Groningen, The Netherlands

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The crystal structure of N<sub>3</sub>P<sub>3</sub>Br<sub>6</sub>, as determined by Giglio & Puliti (1967), has been refined from threedimensional intensity data, collected with Mo radiation on an automatic three-circle diffractometer. Two crystals were used. Data processing and refinement were carried out independently with the data obtained from each crystal. The intensities were corrected for absorption and during the refinement a correction for extinction was applied according to Zachariasen (1967, 1968). Atomic coordinates obtained from the two crystals agree within the standard deviations. The molecules N<sub>3</sub>P<sub>3</sub>Br<sub>6</sub> lie on crystallographic mirror planes of the space group *Pnma*, with unit-cell dimensions a=14.463 (0.002), b=13.410 (0.003) and c=6.601 (0.001) Å. There are four short intermolecular Br  $\cdots$  Br distances (3.65-3.85 Å) and a short N $\cdots$ Br distance (3.24 Å). The two independent angles P–N–P of a molecule are significantly different: 119·3 (0.6) and 122·4 (0.5)°. Other chemically equivalent bond lengths and angles are equal within experimental error. Mean values (and individual standard deviations) are P–N 1.576 (0.008), P–Br 2.162 (0.004) Å, N–P–N 118·5 (0.5) and Br–P–Br 102·1 (0·1)°. The (N–P)<sub>3</sub> ring is non-planar and slightly chair-shaped. The dihedral angles of the ring bonds are -6.7, 6·3 and  $-5.5^{\circ}$ and the largest distance from a ring atom to the least-squares plane through the ring is 0.047 Å.

#### Introduction

The crystal structures of several chloro- and fluorocyclophosphazenes are known with some precision, viz.  $N_3P_3Cl_6$  (Wilson & Carroll, 1960),  $N_4P_4Cl_8$  [two crystal modifications, K form (Hazekamp, Migchelsen & Vos, 1962) and T form (Wagner & Vos, 1968)],  $N_5P_5Cl_{10}$  (Schlueter & Jacobson, 1968),  $N_3P_3F_6$  (Dougill, 1963) and  $N_4P_4F_8$  (McGeachin & Tromans, 1961).

<sup>\*</sup> Part VI: Olthof (1969).

As to the bromo compounds the only crystal structure known is that of the trimer  $N_3P_3Br_6$ . This structure was determined by Giglio & Puliti (1967) from the intensities of 538 X-ray reflexions, estimated visually on Weissenberg films. The accuracy of the results is not high, *e.g.* the values found for the P–Br bond lengths range from 2.14 to 2.20 Å with an individual standard deviation of 0.015 Å.

We intend to investigate the crystal structures of the homologous bromocyclophosphazenes in order to elucidate the conformations of the phosphazene rings. Refinement of the crystal structure of the trimer  $N_3P_3Br_6$ , from new experimental data, seemed necessary, if reliable comparisons with the structures of related compounds are to be made.

Since absorption effects of X-rays in crystals of the bromocyclophosphazenes will be very substantial (linear absorption coefficient for Mo radiation more than 200 cm<sup>-1</sup>), we thought it worth while to use two independent sets of intensities, obtained from two different crystals, for the refinement of the crystal structure of  $N_3P_3Br_6$ , in order to see whether the determined atomic coordinates would be essentially the same.

The results of the refinement of  $N_3P_3Br_6$  are given in the present paper. The crystal structure of the tetramer  $N_4P_4Br_8$ , which has been determined by us, will be described in a following paper.

# **Experimental**

The substance was prepared from phosphorus(III) bromide, ammonium bromide and bromine according to the procedure of Coxon, Sowerby & Tranter (1965). From the reaction mixture, containing trimeric and tetrameric bromocyclophosphazene and possibly higher homologues, suitable crystals of  $N_3P_3Br_6$  were obtained by fractional crystallization from petroleum.

The crystals are orthorhombic. Weissenberg photographs showed the systematic absences 0kl for k+l=2n+1 and hk0 for h=2n+1. The space group therefore is *Pnma* (no. 62) or *Pn2*<sub>1</sub>*a* (no. 33). The unit cell contains four molecules N<sub>3</sub>P<sub>3</sub>Br<sub>6</sub>. In the refinement it was assumed that *Pnma* is the correct space group, implying that the molecules have a crystallographic mirror plane. Accurate unit-cell dimensions were determined from zero-level Weissenberg photographs around the b and c axes. Reflexions from a NaCl crystal were superposed on the films for calibration. The lengths of the unit-cell edges are:

Present work	Giglio & Puliti (1967)
a = 14.463, s.d. $0.002$	Å 14·43, s.d. 0·02 Å
b = 13.410 0.003	13.36 0.02
c = 6.601  0.001	6.63 0.01

Intensities were measured at room temperature on a single-crystal three-circle automatic Nonius diffractometer by the  $\theta$ -2 $\theta$  scan method. All 1176 independent reflexions up to sin  $\theta/\lambda = 0.60$  Å<sup>-1</sup> were measured using Zr filtered Mo radiation. The intensities of all reflexions were measured twice; two different crystals (I and II) of approximate dimensions  $0.1 \times 0.1 \times 0.2$  mm and  $0.2 \times 0.2 \times 0.3$  mm respectively were used. Crystal I gave 995 non-zero reflexions and crystal II 1004. The data for each crystal were processed separately and the refinement was carried out with each series of  $F_{\theta}$  values independently.

The intensities were corrected for Lorentz and polarization effects and for absorption. The latter corrections were calculated with a computer program based on the Busing & Levy (1957) scheme. The linear absorption coefficient for Mo radiation is very high, *viz.* 203 cm<sup>-1</sup>. The transmission factor (1/V)  $\int \exp(-\mu d) dV$  ranged from 0.10 to 0.14 for I and from 0.02 to 0.06 for II. After the absorption correction the  $F_o$  values of the two crystals were compared. After scaling the factor  $R' = \sum ||F_{II}| - |F_{II}|| / \sum |F_{I}|$  was 10.6% for the 899 common reflexions. For the stronger reflexions the scaled  $F_o$  values of II were generally smaller than those of I. As will be seen later this behaviour can only partly be ascribed to extinction effects.

#### Refinement

Least-squares refinements were carried out based on the  $F_o$  values of each crystal separately. A leastsquares program working according to the blockdiagonal approximation (Cruickshank, 1961) was used. The atomic coordinates and thermal parameters from the paper by Giglio & Puliti (1967) were taken as a

Table 1. Final	fractional	coordinates
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Standard deviations are in parentheses.

		Crystal I			Crystal II	
	x	У	z	x	У	Z
N(1)	0.1430 (10)	0.25	0.5967 (30)	0.1434 (11)	0.25	0.5940 (29)
N(2)	-0.0077(7)	0.1499 (8)	0.4566 (19)	-0.0081 (8)	0.1507 (9)	0.4528(24)
P(1)	-0.0605(3)	0.25	0.4121 (9)	-0.0607(4)	0.25	0.4117 (9)
P(2)	0.0907 (2)	0.1477 (3)	0.5556 (6)	0·0910 (3)	0.1480(3)	0.5554(6)
Br(1)	-0.1895(2)	0.25	0.5783 (4)	-0.1895(2)	0.25	0.5789 (4)
Br(2)	-0.1078(2)	0.25	0.1015 (4)	-0.1078(2)	0.25	0.1019 (4)
Br(3)	0.0823(1)	0.0714 (1)	0.8445 (3)	0·0824 (1)	0.0713(1)	0.8442(3)
Br(4)	0.1806 (1)	0.0507 (1)	0·3845 (3)	0·1806 (1)	0.0505 (1)	0.3844(3)

starting point. Atomic scattering factors were taken from Doyle & Turner (1968).

In the initial cycles the scale factor, the atomic coordinates and anisotropic thermal parameters were refined without accounting for anomalous scattering effects and with uniform weights. Later the anomalous scattering by phosphorus and bromine (International Tables for X-ray Crystallography, 1962) was incorporated and a weighting scheme of the form w = $\{w_c^{-1}+p|F_o|^2\}^{-1}$  was introduced. The term  $w_c$  is the weight from counting statistics and p is a constant, chosen so as to make  $(\Delta F)^2$ , averaged over groups of reflexions, as far as possible independent on  $|F_0|$ . The value  $4 \times 10^{-4}$  was given to p for crystal I and  $9 \times 10^{-4}$ for II. Furthermore, the extinction parameter  $p_{ex}$  was refined according to Zachariasen's (1967, 1968) method. This parameter occurs in the function M, which is to be minimized, as follows:

$$M = \sum w[|F_o| - K|F_c| \{1 + C(\theta)p_{\text{ex}}|F_c|^2\}^{-1/4}]^2$$
  
with  $C(\theta) = 10^{-2} (p_2/p_1) (1/\sin 2\theta) \bar{T}$   
and  $p_{\text{ex}} = 2 \times 10^2 (e^2/mc^2)^2 (\lambda^2/V^2) r^*$ 

(See Zachariasen (1968) for the meaning of the symbols.)

At the end of the refinement the indices  $R = \{\sum |\Delta F|^2 / \sum |F_o|^2\}^{1/2}$  and  $R_w = \{\sum w |\Delta F|^2 / \sum w |F_o|^2\}^{1/2}$  were 9.6 and 6.5% for crystal I and 9.4 and 8.0% for II. The theoretical values for R and  $R_w$  based on counting statistics are lower, viz. 8.5 and 4.2% for I, and 7.6 and 5.2% for II. The differences between  $F_o$  and  $F_c$  are thus larger than could be expected on the basis of counting statistics alone, indicating that there are additional errors in the  $F_o$  values. Presumably these errors are mainly due to the difficulties encountered in calculating the absorption corrections. The refined

value of  $p_{ex}$  is 0.005 for I and 0.014 for II. With these numbers the extinction corrections  $y^{1/2} = \{1 + C(\theta)p_{ex}|F_c|^2\}^{-1/4}$  are calculated as about 0.95 and 0.85 for the strong reflexions of I and II respectively. It thus appeared that the extinction effects are rather small in both crystals, though somewhat larger in the bigger one.

The last four cycles of the least-squares refinement were also run without applying extinction corrections. It appeared that this brought about only slight changes in the atomic coordinates, in no case larger than the standard deviations. Furthermore, neglect of the extinction corrections caused only moderate changes in the thermal parameters, the scale factor, the standard deviations and the indices R and  $R_w$ . The sense of these changes is as expected: compared with the refinement with extinction corrections the thermal parameters and scale factor decreased, the standard deviations in all parameters increased and R increased (1% for I and 4% for II). The change in  $R_w$  was less than 1%.

### **Results and discussion**

The final coordinates with their standard deviations calculated by the least-squares program are given in Table 1. It is seen that corresponding coordinates, determined from crystal I and crystal II, are in most cases equal within the standard deviations. The anisotropic temperature factor parameters and their standard deviations are given in Table 2. Here again corresponding values obtained for the two crystals agree within experimental error, indicating that the two sets of  $F_o$  values (after correction for extinction) do not show significant differences in their variation with  $\sin \theta/\lambda$ . Observed and calculated structure factors are compared in Table 3. In this Table the extinction correction is applied to the observed amplitudes  $F_o$ .

Table 2. Parameters (Å<sup>2</sup>×10<sup>-4</sup>) of the anisotropic temperature factors  $\exp \left[-2\pi^2 (h^2 a^{*2} U_{11} + \ldots + 2hka^* b^* U_{12} + \ldots)\right]$ 

Standard deviations are in parentheses.

			Crystal I			
	$U_{11}$	$U_{22}$	U <sub>33</sub>	$2U_{12}$	$2U_{23}$	$2U_{13}$
N(1)	257 (94)	345 (99)	929 (160)	0	0	- 526 (217)
N(2)	349 (65)	262 (64)	652 (90)	-84 (110)	107 (132)	-262(131)
P(1)	222 (27)	386 (31)	439 (34)	0	0 `	- 98 (54)
P(2)	268 (19)	290 (20)	496 (25)	64 (34)	1 (38)	-61(36)
Br(1)	375 (12)	615 (15)	620 (16)	0	0	233 (26)
Br(2)	555 (15)	752 (17)	441 (14)	0	0	- 224 (26)
Br(3)	669 (12)	655 (11)	519 (10)	-6 (19)	274 (19)	9 (19)
Br(4)	509 (9)	514 (9)	702 (12)	197 (17)	-216 (20)	271 (19)
			Crystal II	[		
N(1)	205 (88)	409 (101)	610 (130)	0	0	- 439 (184)
N(2)	326 (70)	337 (72)	952 (121)	-88(123)	-168 (165)	- 234 (156)
P(Ì)	201 (26)	389 (30)	458 (34)	0`´	0` ´	-131 (52)
P(2)	274 (19)	278 (19)	462 (23)	48 (35)	18 (38)	- 79 (35)
Br(1)	381 (13)	592 (15)	608 (16)	0	0	208 (25)
Br(2)	587 (16)	755 (18)	408 (14)	0	0	-155(25)
Br(3)	692 (12)	629 (11)	526 (10)	9 (20)	289 (20)	38 (19)
Br(4)	506 (10)	513 (10)	728 (12)	230 (17)	- 196 (21)	238 (19)

is given in Fig. 1. The  $(N-P)_3$  rings of the molecules A and C lie approximately at height z=0.5, and those corresponds to that used in the Tables. Fig. 2 shows a

A projection of the crystal structure along the c axis of the molecules B and D approximately at height z=0. The numbering of the atoms in molecule A

# Crystal I

L FO FL AC	80 M K.	- FO FC AC	вс н к.	. FC FC	AC 6C	 10 FC	. AC 84	 +0 +0	C AC	вс н	K L	F0 FC	A C	ьC
		1   2.7.9   4.3.9   2.9.9     1   1.7.9   4.3.9   2.9.9     1   1.1.9   1.9.9   3.9.9     1   1.1.9   1.9.9   3.9.9     1   1.2.4.4   3.2.9   2.3.1.5     1   2.2.4.4   3.2.9   2.3.1.5     1   2.2.4.0   2.2.9   2.2.1     1   2.2.4.0   2.2.9   -2.2.1     1   2.2.4.0   2.2.9   -2.2.1     1   2.9.7   3.2.2   -3.9.8     1   2.9.7   3.2.2   -3.9.8     1   2.9.7   3.2.2   -3.9.8     1   2.9.7   3.2.2   -3.9.8     1   2.9.7   3.2.2   -3.9.8     1   2.9.7   3.2.2   -3.9.8     1   2.9.7   3.2.2   -3.9.8     1   2.9.7   3.2.2   -3.9.8     1   2.9.7   3.2.2   -2.9.1     1   2.9.7   3.9.3   3.2.2 </td <td></td> <td></td> <td></td> <td>257 (2000) 257 (2</td> <td></td> <td>2011 13 13 13 13 13 13 13 13 13 13 13 13 1</td> <td>33 -</td> <td></td> <td></td> <td></td> <td></td> <td>329943111-144511122217104511288319977284529977284592728459284392250344228413737331121431980453333311124319214454284519805941241242845198059444552895194545111245459919214454528951945454511111111111111111111111111111</td>				257 (2000) 257 (2		2011 13 13 13 13 13 13 13 13 13 13 13 13 1	33 -					329943111-144511122217104511288319977284529977284592728459284392250344228413737331121431980453333311124319214454284519805941241242845198059444552895194545111245459919214454528951945454511111111111111111111111111111

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3	85	234	264	262	34	0	0.6	356	420	-419	- 25	1	36	559	557	-552	-72	8		105	25	- 25	,	1	17	60	42	41	10	- Å	57	150	152	-147	-37
- 4	85	856	839	833	107	1	0 8	782	756	-751	-89	ž	3 6	245	155	154		9	6 6	251	87	-87	-22	2	17	198	124	123	12	1	67	148	190	-187	-32
5	85	250	15/	155	54	ž	0.6	367	399	393	69	3	3 5	452	450	446	56	2	76	388	423	420	53	3	17	346	333	-330	-45	2	67	236	105	105	3
	85	400	387	- 384	-50	3	06	474	579	-577	-57	4	3 5	419	477	-469	-84	3	76	249	314	312	41	4	17	323	298	-296	-36	0	77	77	85	83	18
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projection along the *b* axis of the molecules lying on the mirror plane y=0.25. All independent intermolecular distances which are nearly equal to or shorter than the sum of the corresponding van der Waals radii are given in the Figures (Pauling, 1960; r(N)=1.5, r(P)=1.9, r(Br)=1.95 Å).

It is interesting to note that in the crystal structure of  $N_3P_3Cl_6$  (Wilson & Carroll, 1960) only the intermolecular distances of the type given in Fig. 2 are shorter than the corresponding van der Waals distances, whereas in the crystal structure of  $N_3P_3F_6$  (Dougill, 1963) no short intermolecular distances occur.

The short distance  $N(1)\cdots Br(1)$ , shown in Fig. 2, is part of the approximately linear array  $N(1)\cdots Br(1)-P(1)$ . Likewise the short distance  $Br(1)\cdots Br(2)$  is part of the linear array  $Br(1)\cdots Br(2)-P(1,B)$ . Such linear arrays have been observed in the crystal structures of polyhalogen compounds and have there been interpreted in terms of three-centre four-electron bonding (Wiebenga &



Fig.1. Projection of the crystal structure along the c axis. The molecules lie on mirror planes at y=0.25 and y=0.75. The rings of the molecules A and C are approximately at height z=0.5, and those of B and D approximately at height z=0. The short intermolecular distances are as follows: Br(3, A)  $\cdots$  Br(3, C[z+1]), 3.68 Å; Br(3, A)  $\cdots$  Br(4, D[z+1]), 3.81 Å; Br(3, A)  $\cdots$  Br(4, A[z+1]), 3.85 Å.

Kracht, 1969). A similar type of interaction may be present in the arrays  $N(1)\cdots Br(1)-P(1)$  and  $Br(1)\cdots Br(2)-P(1,B)$  of Fig. 2. It may be noted that these arrays are approximately perpendicular to each other; such a perpendicular orientation has also been observed with the polyhalogens.

As seen in Fig. 1, there is a short intermolecular distance  $Br(3, A) \cdots Br(3, C[z+1])$ , which is comparable in length to  $Br(1) \cdots Br(2)$ , Fig. 2. However,  $Br(3, A) \cdots Br(3, C[z+1])$  is not part of a linear array of the type just considered: the angle P(2, A)-Br(3, A)-

Br(3, C[z+1]) is 141°. Also, in the crystal structure of N<sub>4</sub>P<sub>4</sub>Br<sub>8</sub> short Br...Br distances (3.66, 3.74 Å) are observed (Zoer & Wagner, to be published), that do not form linear arrays with a phosphorus atom.

The molecule N<sub>3</sub>P<sub>3</sub>Br<sub>6</sub> is shown in Fig. 3 and the bond lengths and valence angles, calculated from the mean atomic coordinates of I and II, are given in Table 4. The standard deviations in this Table were derived from the standard deviations in the mean coordinates  $\sigma_m = \frac{1}{2} \{\sigma^2(I) + \sigma^2(II)\}^{1/2}$ . For comparison Giglio & Puliti's results have been added to Table 4.

# Table 3 (cont.)

# Crystal II

Table 3 (cont.)

Crystal II

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	11 2	24	238 222	-221	-17	10 7	342	316	-314	-40	7	0 5	440	432	-430	-44	10	55	205	153	-153	-7	10	0.6	299	240	-237	-40	5 7 5	60	17	-15	.,
	13 2	4	97 18	-183	-23	12 7	311	325	321	56	ş	ŏŝ	763	788	-782	-95	12	55	2223	25/	-256	-17	12	16	197	135	-135	-4	675	187	161	160	16
	1 4	34	483 414	-412	+20	0 8	392	453	448	66	11	05	417	391	386	64	1	0 5	257	23	-21	10		1.6	127	158	-158	-15	0 0 6	769	759	-753	-96
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6 0 4 453 459 -266 47 286 185 4 52 154 -52 154 -53 5 4 5 4 5 20 104 104 10 30 011 5 126 5 3 17 13 7 5 6 168 29 24 17 6 47 286 185 183 30 7 8 4 112 26 17 -22 4 12 146 138 25 7 4 5 131 64 10 30 111 5 116 30 111 5 126 19 26 143 11 0 57 516 526 519 66 7 8 4 127 397 -354 -457 512 4 99 116 -115 -11 8 4 5 22 190 15 24 99 24 10 9 5 6 433 413 318 44 1 5 7 66 133 152 14 7 8 4 15 54 97 21 21 2 6 12 4 99 116 -115 -11 8 4 5 22 190 15 51 5 24 99 24 9 5 6 433 413 318 44 1 5 7 66 133 152 14 10 8 4 65 40 5 40 -51 24 99 116 -11 9 10 42 42 4 10 8 4 65 40 5 40 5 43 64 133 64 113 4 84 159 199 9 12 4 5 247 244 -23 -23 21 25 218 218 -212 -41 1 2 6 6 20 129 -20 -23 4 5 7 129 144 -12 1 2 16 4 -159 -30 11 8 4 55 50 50 -9 -9 2 13 4 307 387 -533 -55 0 5 3 356 361 361 -361 -361 -361 -361 -361 -361	4 0		654 650	-055	-//	0 12	4 178	224	-223	-24	3	4 5	275	225	-553	-31	7.1	0 5	336	252	250	31	5	5 6	551	563	-580	-59	5 4 7	355	331	328	43
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4 7 4 565 48V -484 -73 3 0 5 533 635 -630 -73 4 5 5 393 307 307 25 5 0 6 222 180 177 30 1 7 6 208 138 -136 -19	3 7		572 603	, aly , ., ., ., ., .,	-71	2 0	5 352 5 310	299	-290	-73	23	55	735	725	-721	- 82	2	06	361 571	396 591	590 880-	69 •58	ŝ	6 6	312	288	266	41				-	
	4 2	7.4	565 481	-484	-75	3 0	5 63	635	-630	-73	1	5.5	393	397	397	25	5	0 6	222	180	177	30	1	7 6	208	138	-136	-19					

The P-Br bonds (mean value 2.162 Å) are equal within experimental error, as well as the angles Br-P-Br (mean  $102\cdot1^{\circ}$ ). Within the (N-P)<sub>3</sub> ring no significant differences are found between the P-N bonds (mean 1.576 Å), or between the angles N-P-N (mean 118.5°). However, there is a small, but significant, difference (3°) between the two independent angles P-N-P (119.3 and 122.4°). The smaller value of the angle P(2)-N(1)-P(2'), as compared with the angle P(2)-N(2)-P(1), together with the differences in the P-N bonds and the angles N-P-N (though the latter differences are not significant) describe a slight deformation of the six-membered ring of a type consistent with the bonding interaction, discussed above, between the nitrogen atom N(1) and the group P(1)-Br(1) of a second molecule

The  $(N-P)_3$  ring is not entirely planar. The dihedral angles assignable to the P-N bonds are:  $-6.7^{\circ}$  for N(1)-P(2),  $6.3^{\circ}$  for P(2)-N(2) and  $-5.5^{\circ}$  for N(2)-P(1), indicating a slight chair-shaped conformation. The best plane through the six ring atoms, the phosphorus atoms having a weight four times that of the nitrogen atoms, is given by the equation

-0.3958 X + 0.9183 Z = 2.836 (X and Z in Å).

The distances from the ring atoms to this plane and their standard deviations are  $\Delta N(1) = -0.047$  (0.011),  $\Delta P(2) = 0.011$  (0.003),  $\Delta N(2) = -0.035$  (0.008) and  $\Delta P(1) = 0.008$  (0.004) Å. The  $\chi^2$  test (Cruickshank & Robertson, 1953) on the planarity of the ring yielded  $\chi^2 = 9990$  and P < 0.001 for n = 6.

In the crystal structures of the related compounds  $N_3P_3Cl_6$ ,  $N_3P_3Cl_5F$  (Olthof, 1969) and  $N_3P_3F_6$  the

 $(N-P)_3$  ring has been found to be planar within experimental error. The departure from planarity in the case of  $N_3P_3Br_6$  is most likely due to intermolecular steric interactions, though the non-planarity is not apparent from the short intermolecular contacts, presented in Fig. 1 and 2.

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Fig.2. Projection along the *b* axis of the molecules lying on the mirror plane y=0.25 and short intermolecular distances in this plane.

	Presen	t work	Giglio & F	uliti (1967)
N(1) - P(2)	1·588 Å	0·006 Å	1·55 Å	0∙035 Å
N(2) - P(1)	1.565	0.008	1.60	0.033
N(2) - P(2)	1.575	0.008	1.60	0.045
P(1) - Br(1)	2.165	0.004	2.20	0.011
P(1) - Br(2)	2.158	0.004	2.17	0.017
P(2) - Br(3)	2.168	0.003	2.20	0.019
P(2) - Br(4)	2.159	0.003	2.14	0.013
		<b>a</b>	10(0)	• • • •
P(2) - N(1) - P(2')	119·3°	0.6°	126.8	2.6
P(2) - N(2) - P(1)	122.4	0.5	118.6	2•4
N(2) - P(1) - N(2')	117.4	0.6	118.3	1.8
N(1) - P(2) - N(2)	119.0	0.4	115.8	2.0
Br(1)-P(1)-Br(2)	102.1	0.1	101.7	0.4
Br(3)-P(2)-Br(4)	102.0	0.1	103.4	0.4
N(1) - P(2) - Br(3)	106.9	0.2	106.7	2.4
N(1) - P(2) - Br(4)	108.7	0.4	111.9	1.7
N(2) - P(2) - Br(3)	109-3	0.4	114.4	1.8
N(2) - P(2) - Br(4)	109.7	0.4	104.0	1.8
N(2) - P(1) - Br(1)	109.1	0.4	113.4	1.7
N(2) - P(1) - Br(2)	109.0	0.4	103.8	2.1

Table 4.	Bond	lengths	and val	'ence angl	es with	ı standar	d deviations



Fig. 3. Molecule N<sub>3</sub>P<sub>3</sub>Br<sub>6</sub>.

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